

ELECTROCONDUCTING POLYMERS AS A CORROSION PROTECTION MATERIALS REVIEW

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INTRODUCTION

Electroconductive polymers (ECP) have been extensively studied during last years. First experiments were carried on noble material supports (like gold, platinum or glassy carbon) but more practical attempts concern ECP deposited onto active metals support. It was found that the ECP can be deposited onto the most of active metals like iron, aluminium, zinc, copper, nickel and their alloys. Polymers can be deposited in both ways: chemically or electrochemically, from organic solvents or water solutions [1] inspite of the simultaneous processes of anodic metal dissolution. However, the dissolution or passivation of active metal as a result of electrosynthesis of polymer film during anodic polarisation is not the only problem connected with using active supports. The further difficulties concern stability of polymer modified support in corrosive conditions of electrolytes used in devices. Among numerous possible applications of the ECPs films in practice the possibility of their application in corrosion protection of metals recently have been studied intensively. Below the papers devoted to this problem which have been published in 2006-2007 years are reviewed.

Most of publications concern with electrosynthesis of the ECPs layer as a protective films for iron [2-8] and mild steel, carbon steel or stainless-steel [9-54], but non-ferrous metals as a support for electrodeposition of ECPs layer have been also used: aluminium and its alloys [55-63], copper [64-73], nickel [74-76], zinc [77,22] and magnesium alloys [78,79].

Among the polymers more often were examined: polypyrrole (PPy) [2, 3, 5, 7, 17, 19-22, 33, 34, 41; 42; 47; 54-59; 68, 70-72, 75, 77] or its composite

e.g. [27] and polyaniline (PANI) [4,9-11, 28, 35, 36, 40, 43, 48, 51; 54; 56; 61; 62; 63, 65, 73, 74], or its composite film like: PANI + TiO₂ [12,14] or PANI with acrylic and epoxy [18] and bilayer PPy/PANI was also recommended as a protective coating [31]. Conducting pyrrole derivatives film for nickel protection against corrosion was studied [76]. Among other polymers poly(N-ethylaniline) (PNEA) [8,16,25,26] and poly(o-anisidine) [13, 66] also attract investigators attention.

Application of the copolymer film has been suggested, e.g. N-methylpyrrole; 3,4-ethylenedioxythiophene [37,38]. Also heterogeneous blends of a soluble substituted polyaniline, poly(2,5-dimethoxy aniline) (PDMA), and two fluoropolymers, poly(vinylidene fluoride) and poly(tetrafluoroethylene-co-vinylidene fluoride-co-propylene) have been prepared and examined on steel [39]. Using sulfonated polypyrrole [44] and hybrid materials [24] as well as electrochemically synthesized polypyrrole coating modified with very thin graphite layer and top coated with another polypyrrole film [47] as a protective coating for steel were proposed. Poly(2,5-dimethylaniline) coatings and poly(o-toluidine) (POT) coatings were synthesized and examined on copper [65,67].

Conductive polymers may be used too as pigments in anticorrosive coatings [80-87].

CONCLUSION

The electroconducting polymers are interesting materials for corrosion protection and a number of publications concerning them still growing.

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